

Build a Binary Ionic Compound

Binary ionic compounds are composed of a metal which has lost electrons (oxidation) and a nonmetal which has gained electrons (reduction).

You will receive two color-coded index cards, one with the name of a metallic element, and the other with the name of a non-metallic element. It will be your responsibility to produce a 2-dimensional artistic model of the formation of an ionic compound from these two elements. Because some of you have more “artistic ability” than others, you have the choice of submitting your finished product as a hard copy (on paper or other media) or as a digital work in Google Classroom.

Expectations:

1. Your project will include the correct names (2 pts) and number of atoms (2 pts) used to form the compound
2. It will include images (your own work) of the atoms you begin with (2 pts), and the ions formed (2 pts)
 - a. Show the correct number of electrons being transferred (2 pts)
 - b. Do metals get bigger or smaller as they lose electrons? (2 pts)
 - c. Will nonmetals get bigger or smaller as they gain electrons? (2 pts)
3. It will include the correct name and formula of the ionic compound produced. (2 pts)
4. Include at least two facts about potential uses of this ionic compound (2 pts)